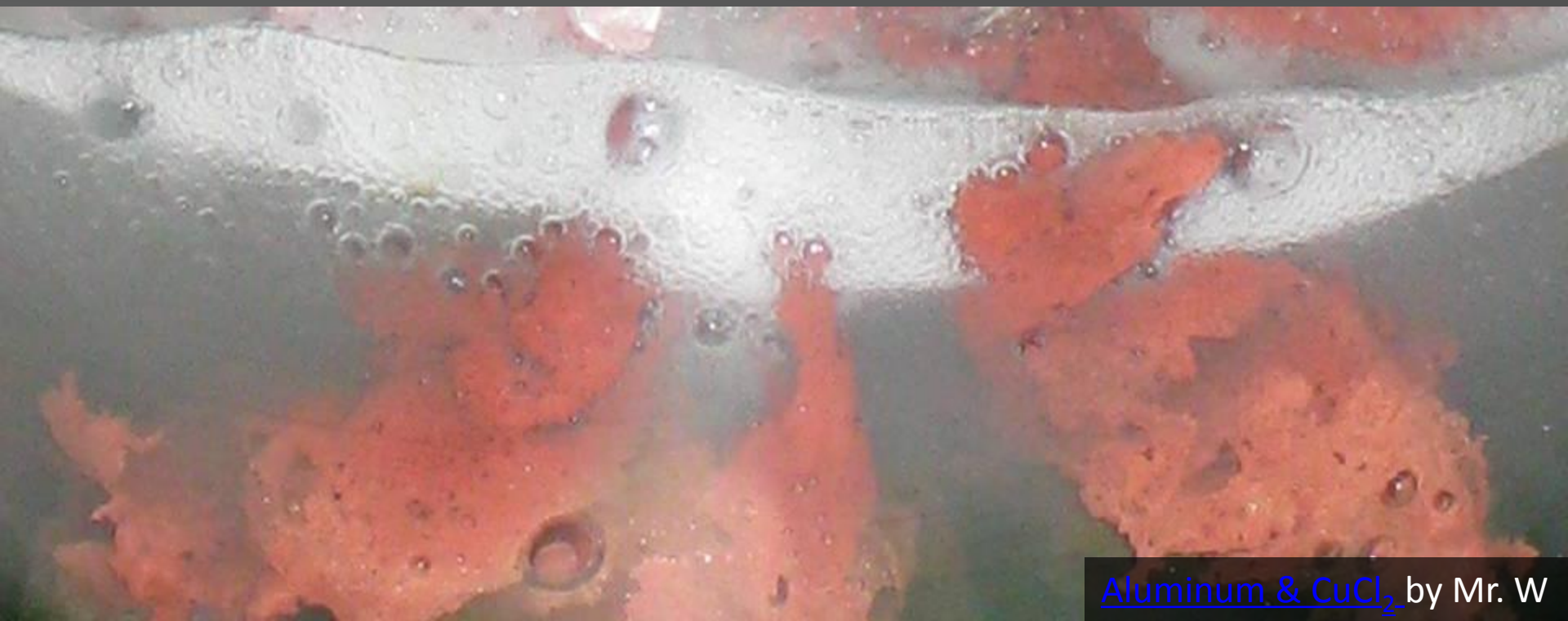
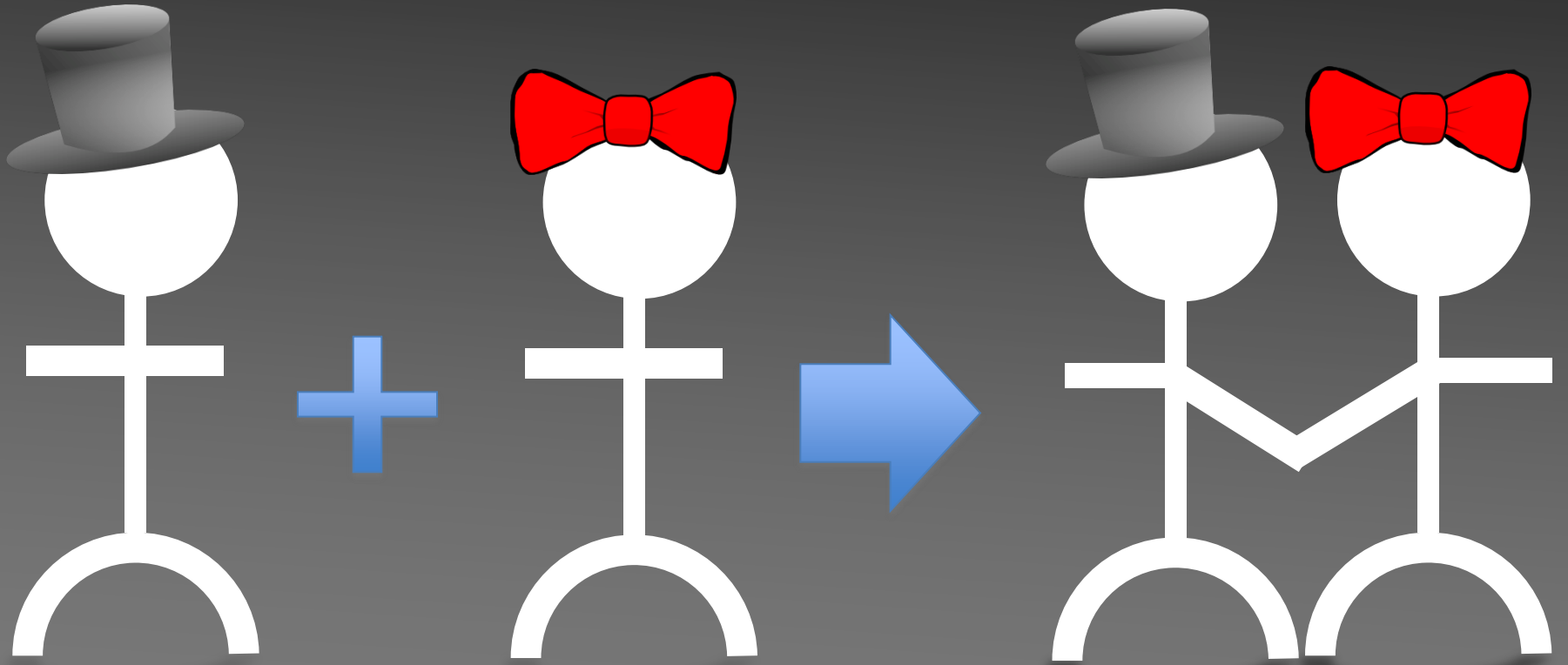


Types of Chemical Reactions



/Synthesis/

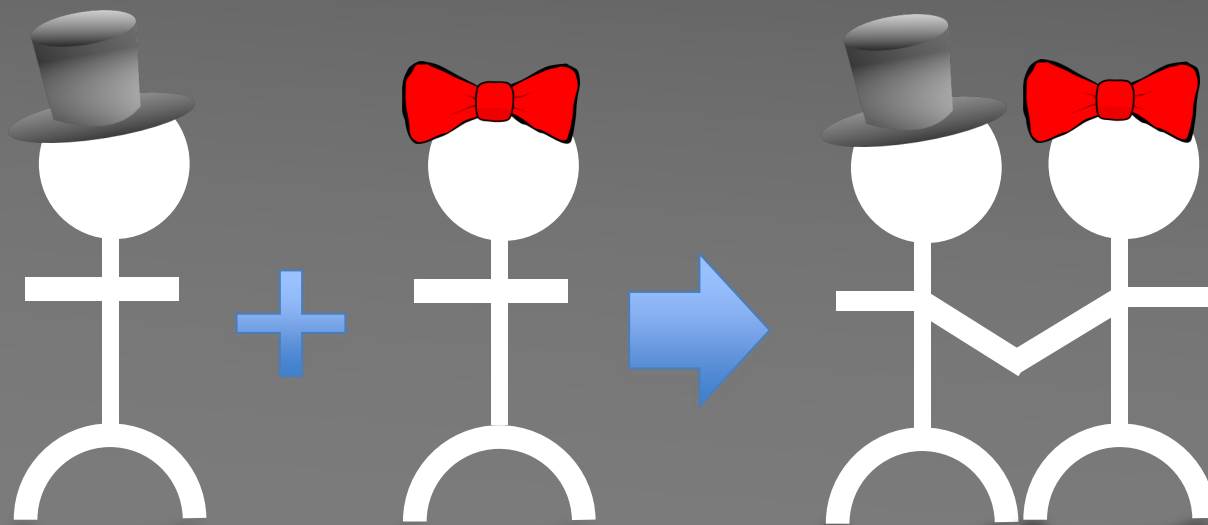
The get-together



/Synthesis/

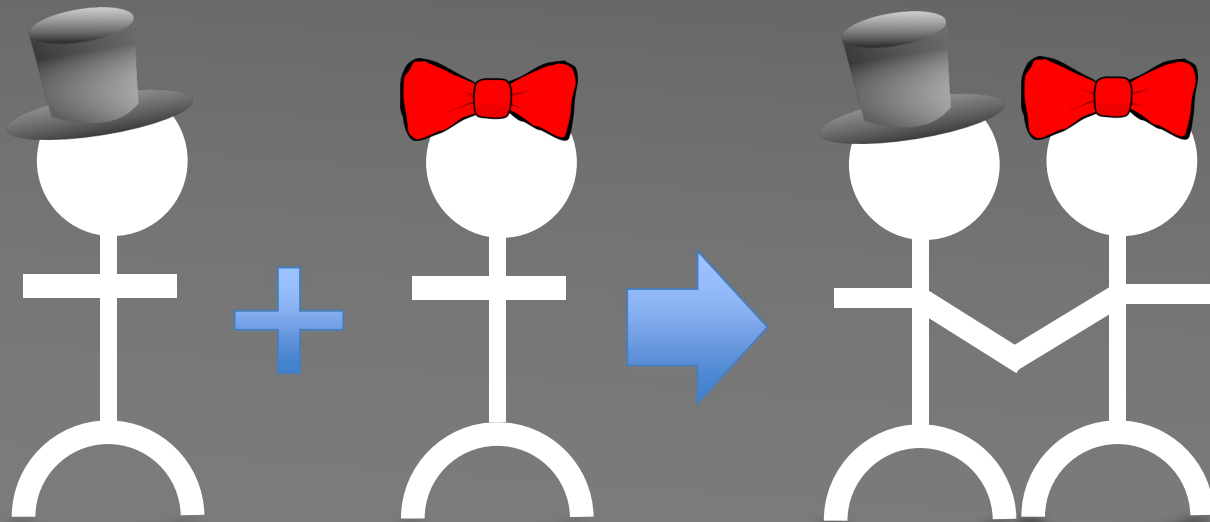
The get-together

Two or more chemicals bond together forming one new substance.



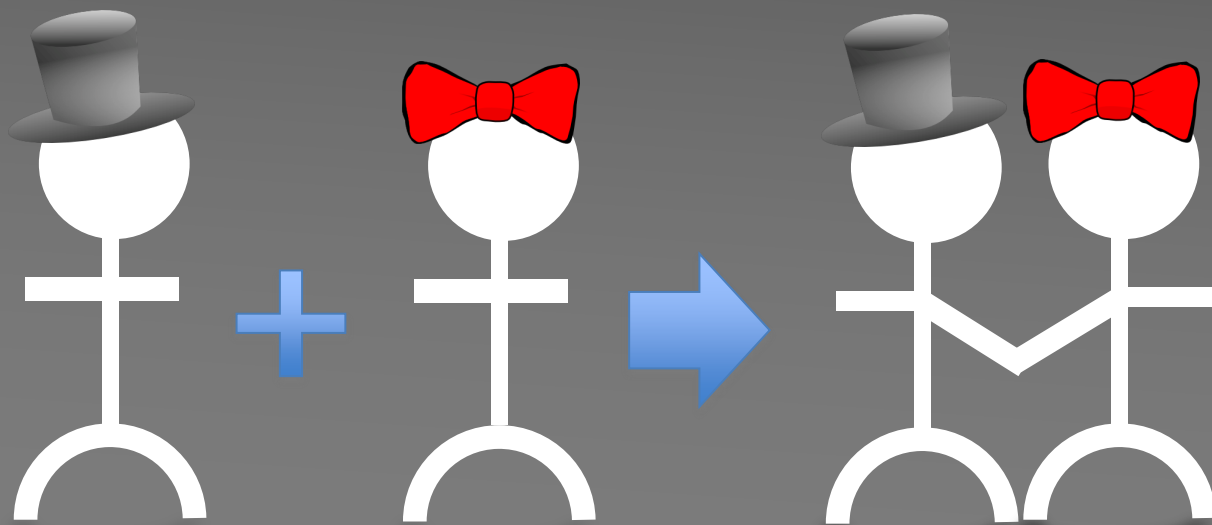
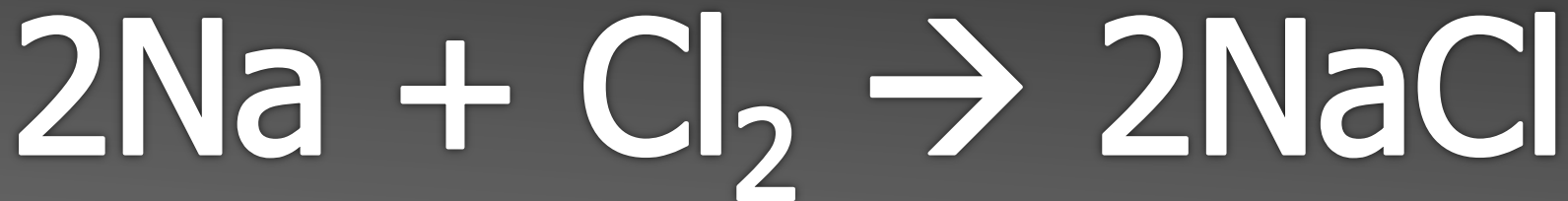
/Synthesis/

The get-together



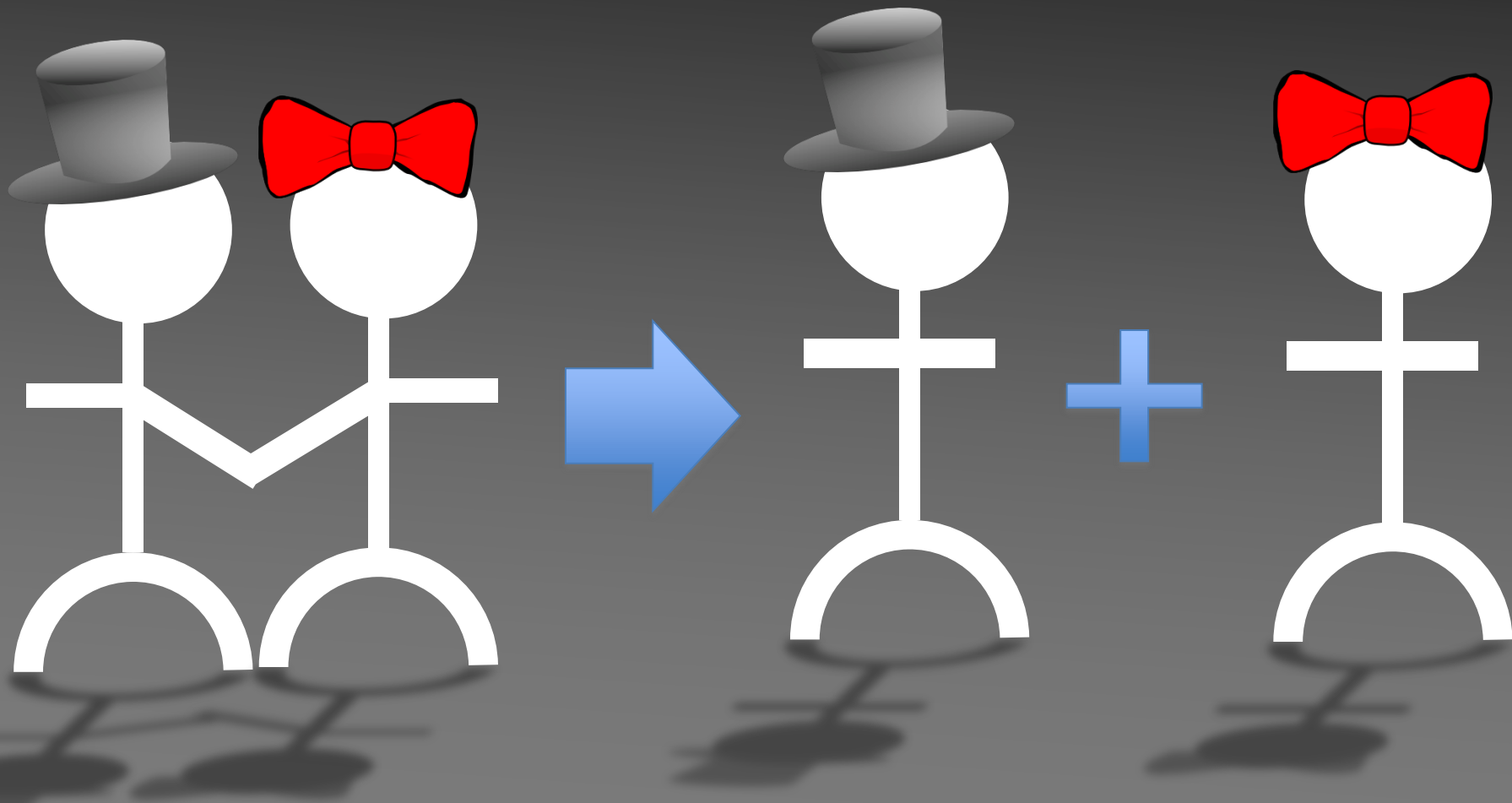
/Synthesis/

The get-together



/Decomposition/

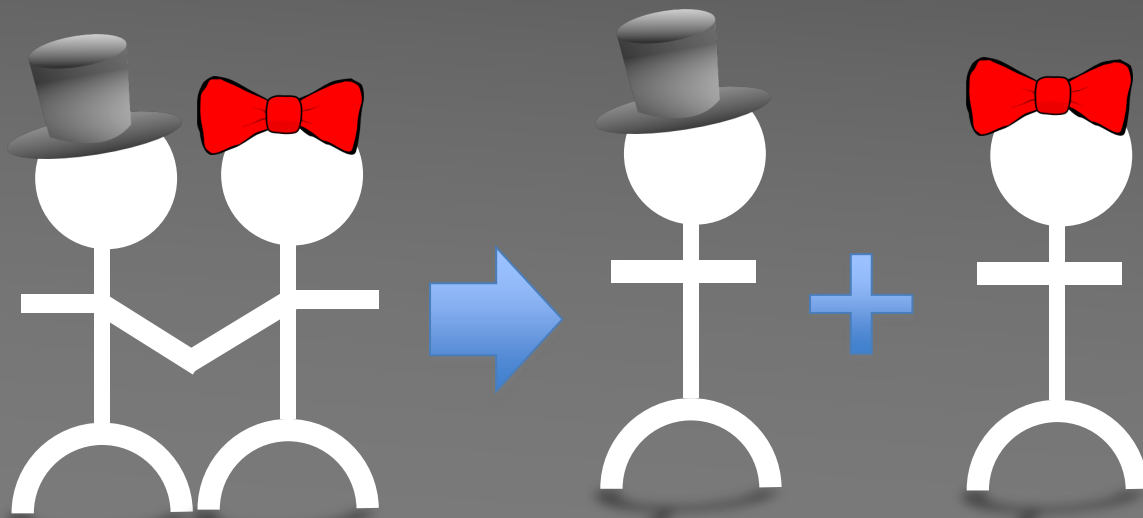
The break-up



/Decomposition/

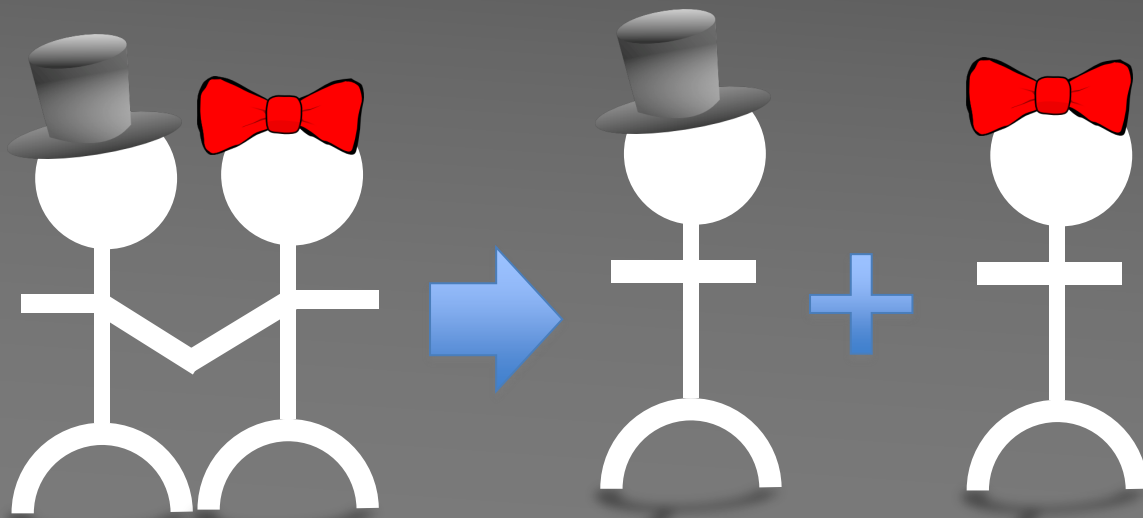
The break-up

One substance breaks down into two or more separate substances.



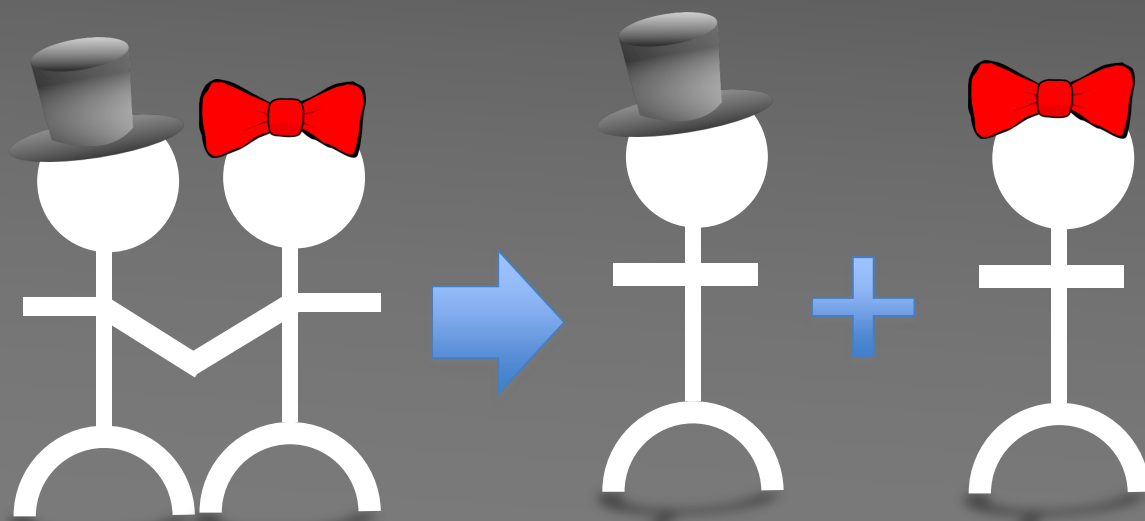
/Decomposition/

The break-up



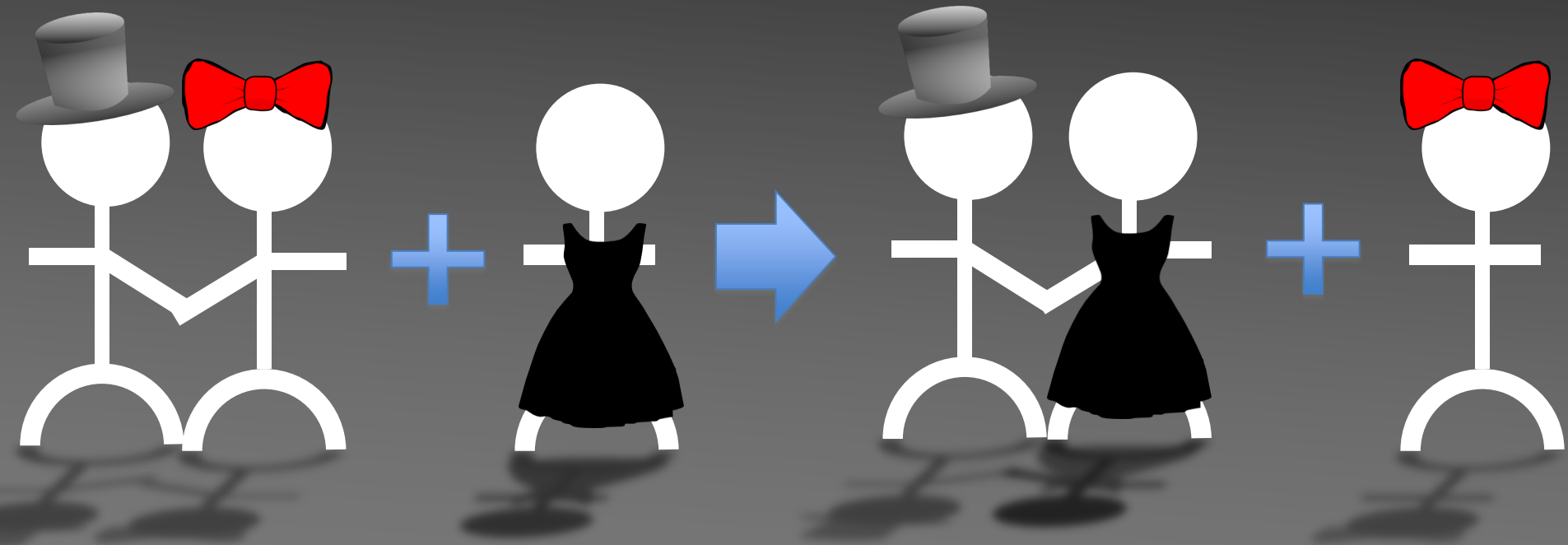
/Decomposition/

The break-up



/Single Replacement/

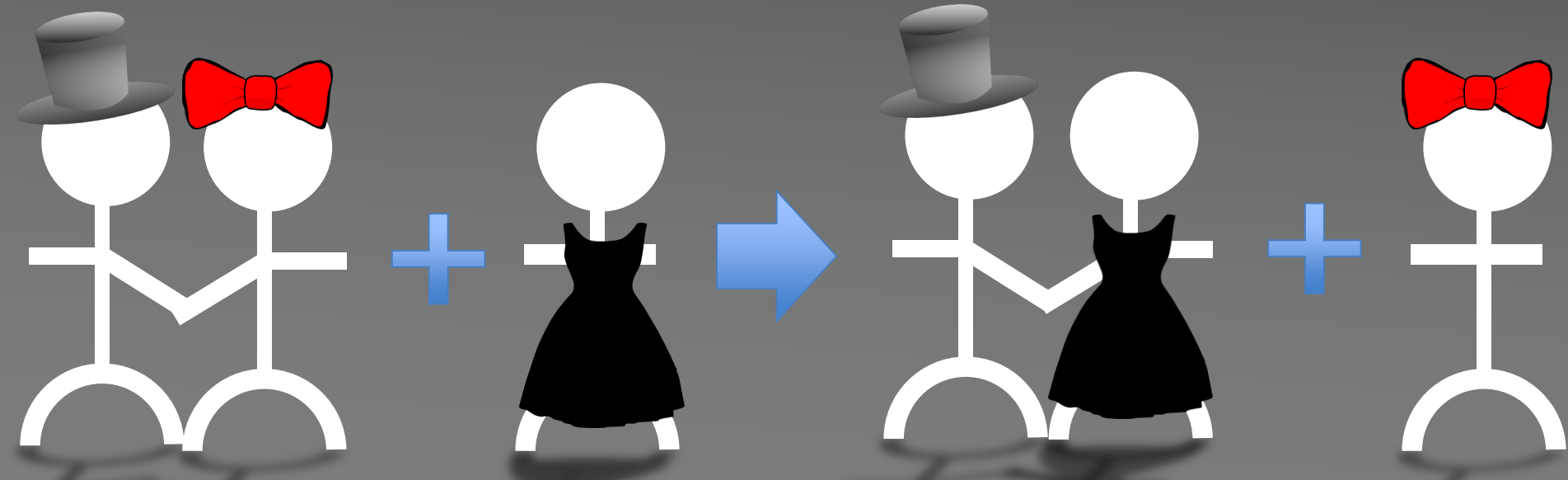
The cheater



/Single Replacement/

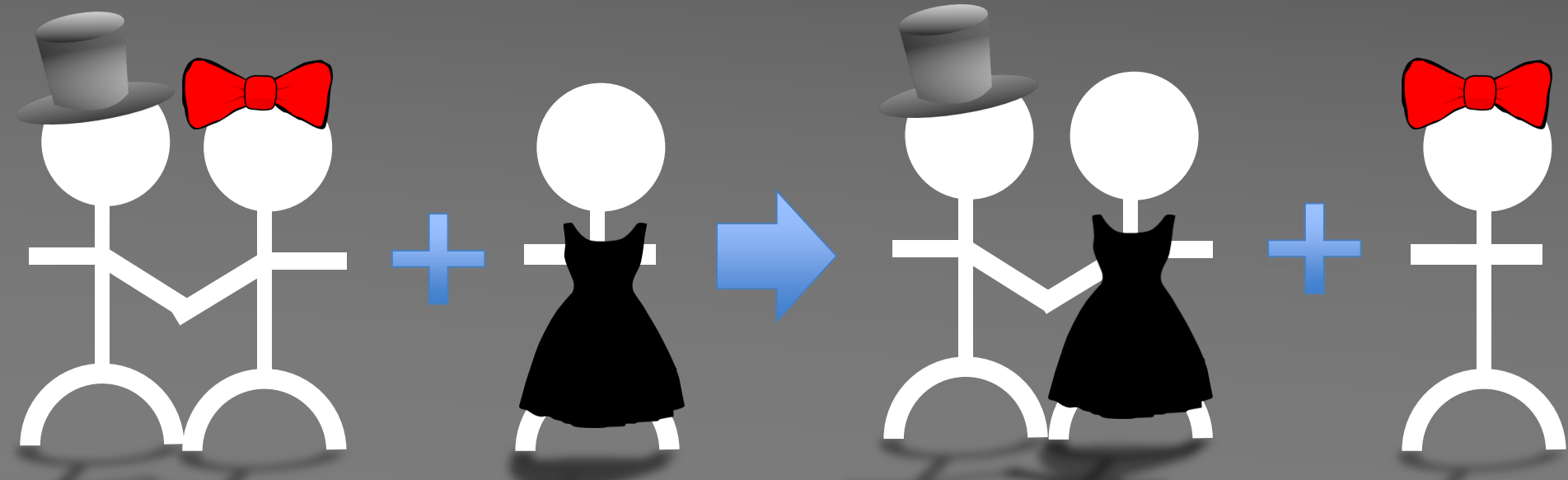
The cheater

One element knocks another element out of a compound.



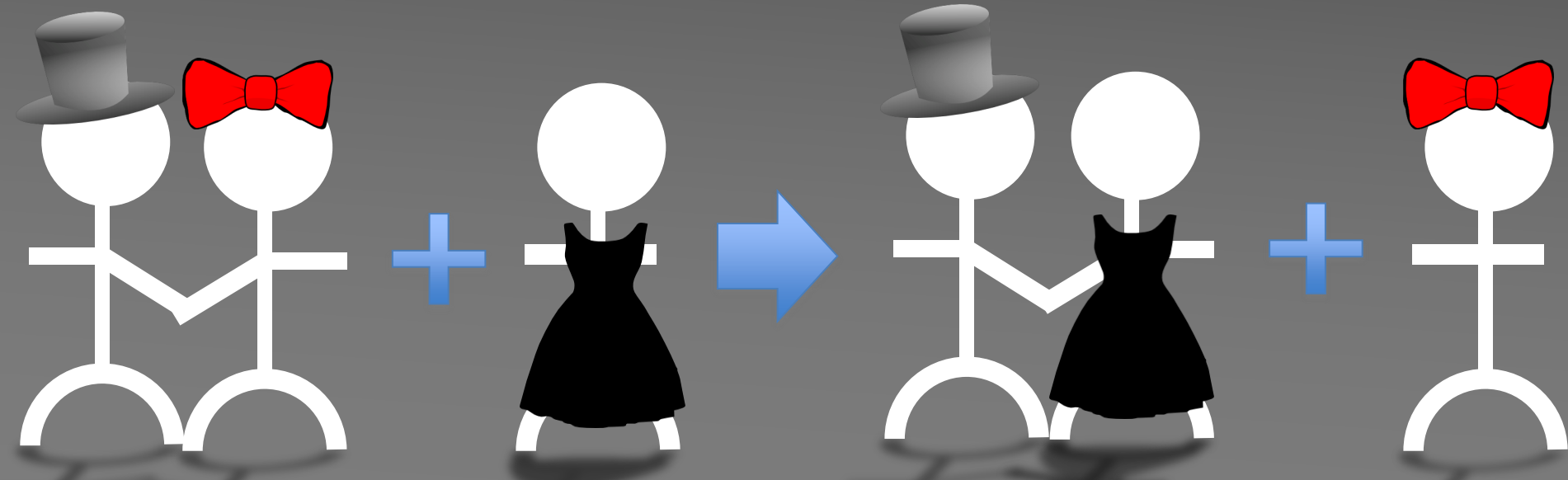
/Single Replacement/

The cheater



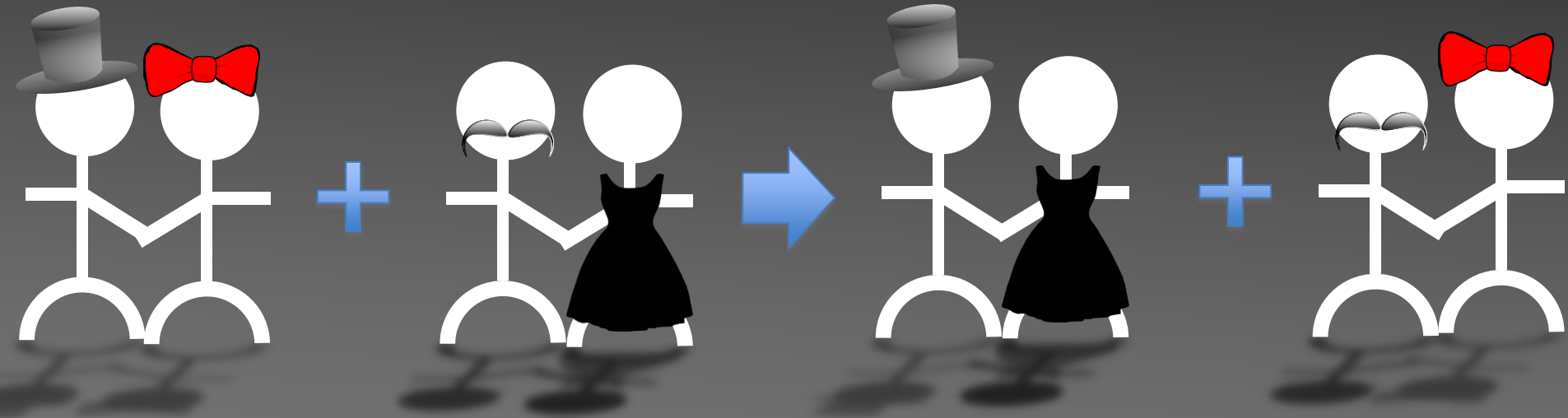
/Single Replacement/

The cheater



/Double Replacement/

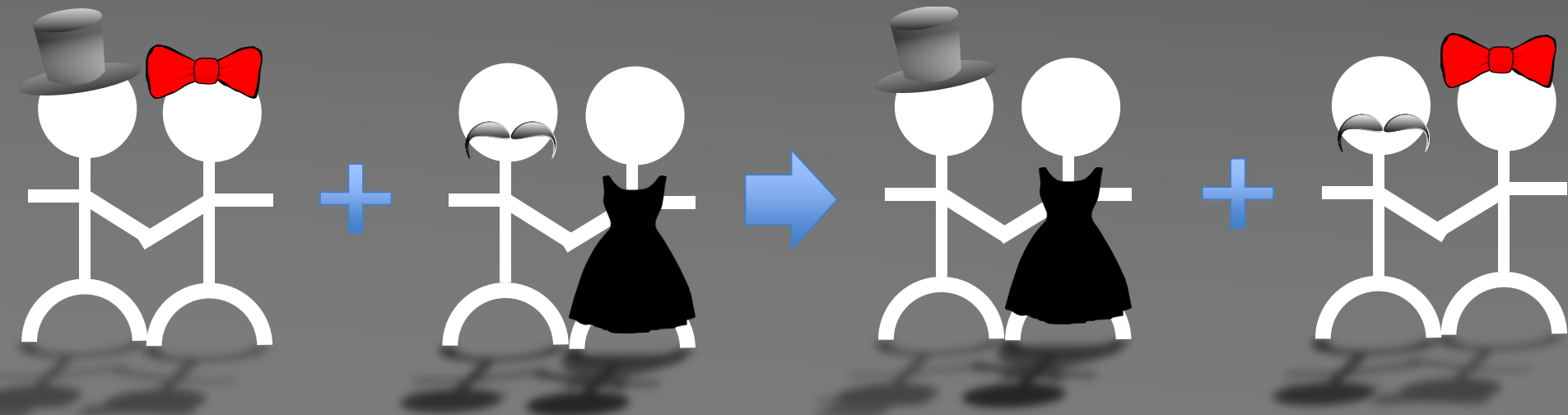
The swap



/Double Replacement/

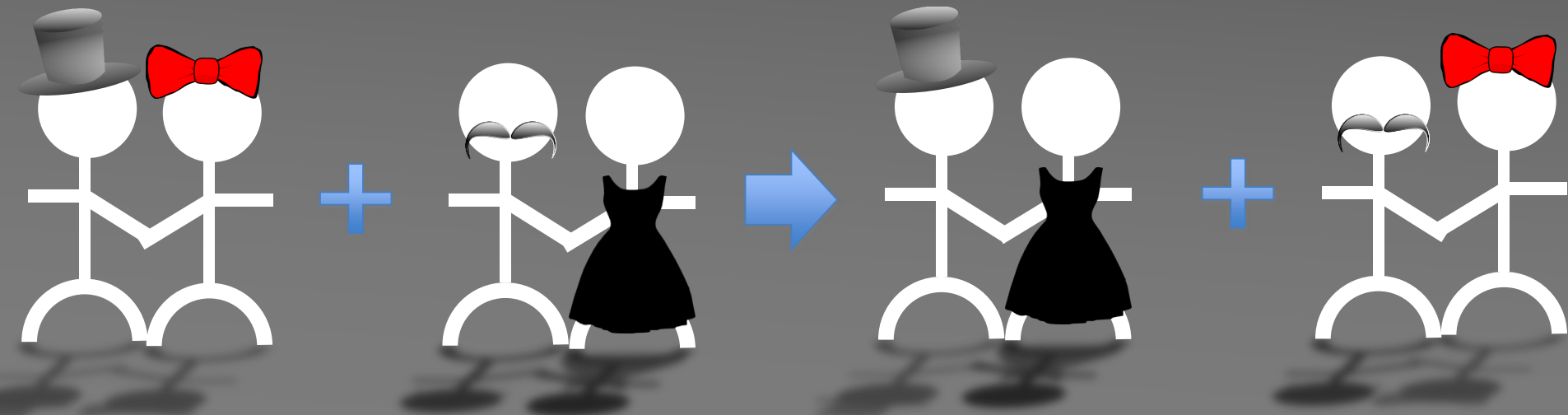
The swap

Two compounds switch ions with each other.



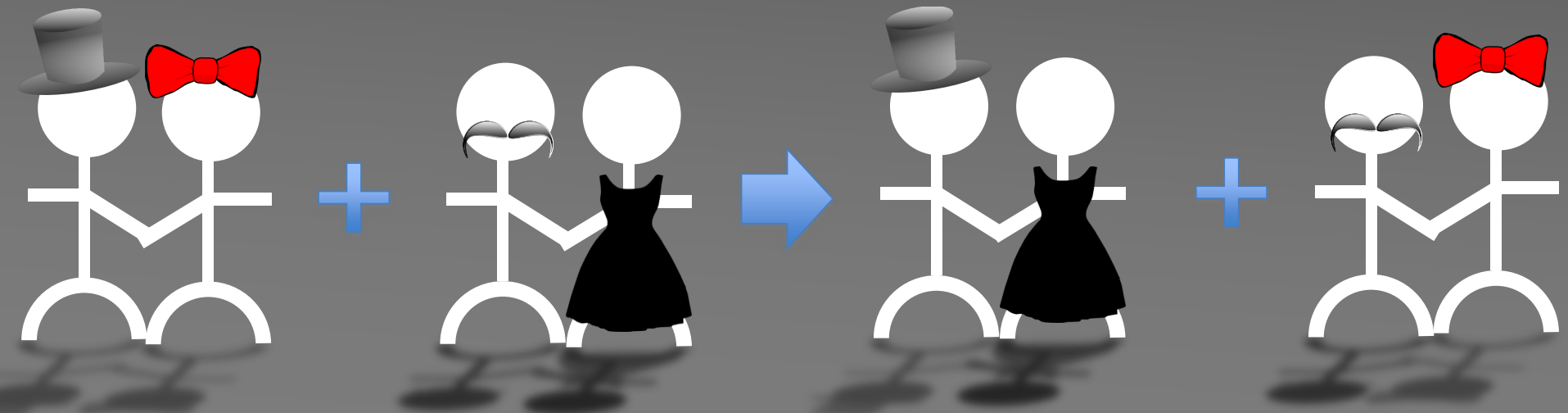
/Double Replacement/

The swap



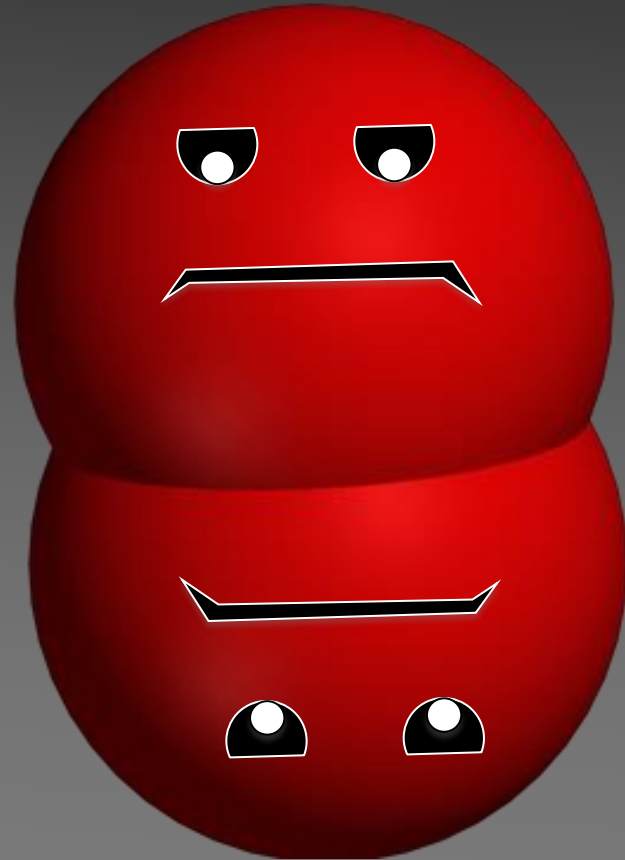
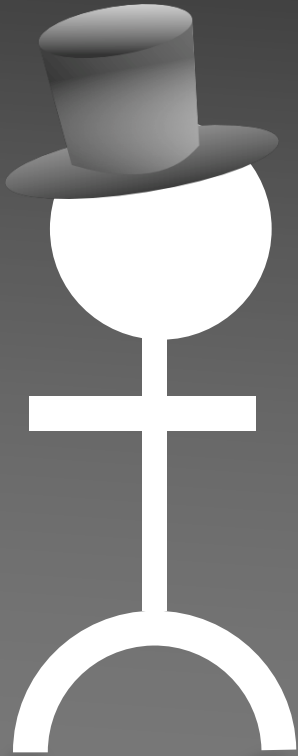
/Double Replacement/

The swap



/Combustion/

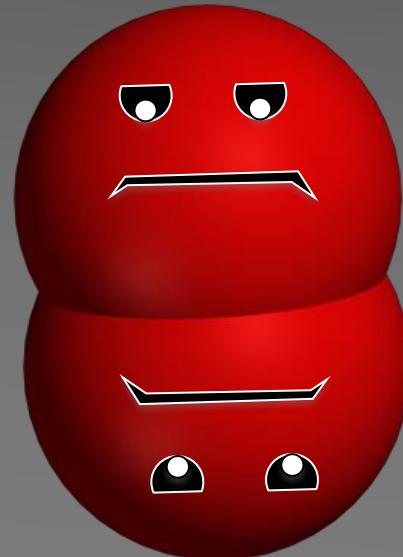
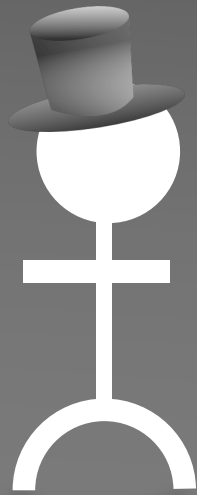
Everyone loves O₂



/Combustion/

Everyone loves O₂

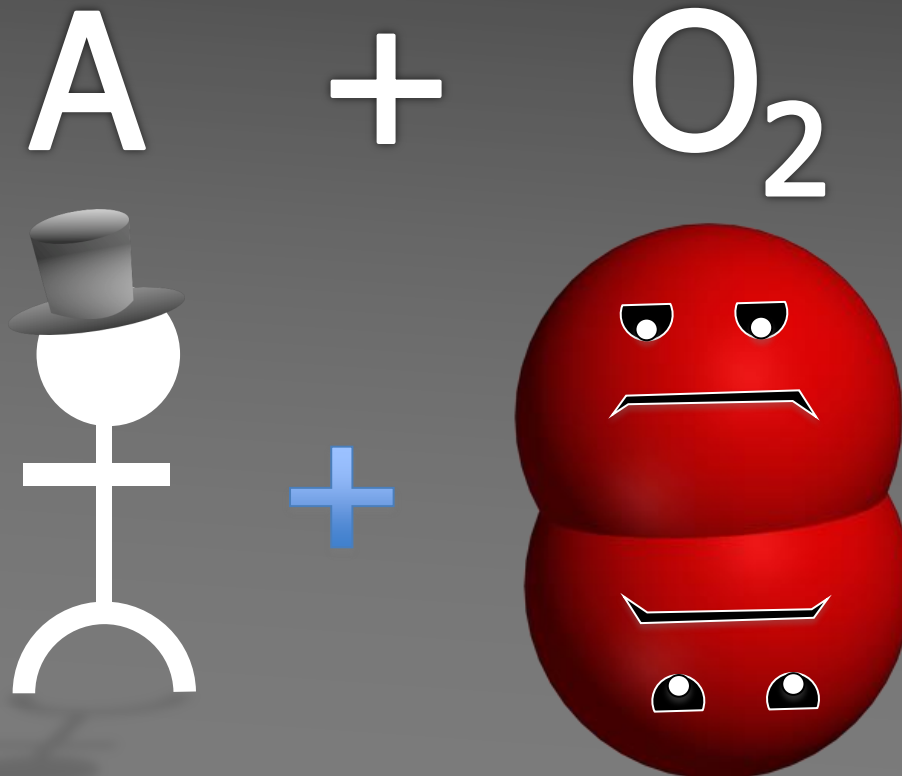
A compound burns in oxygen gas.



/Combustion/

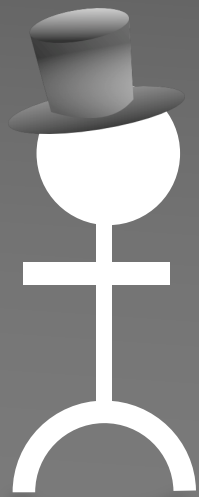
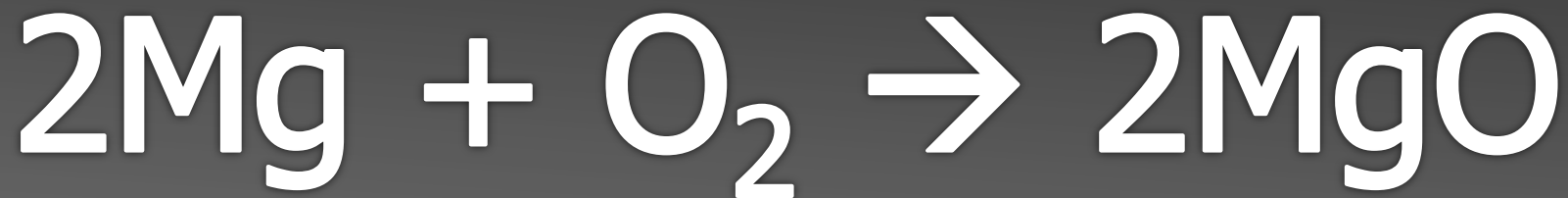
Everyone loves O₂

Oxygen will always be a reactant.

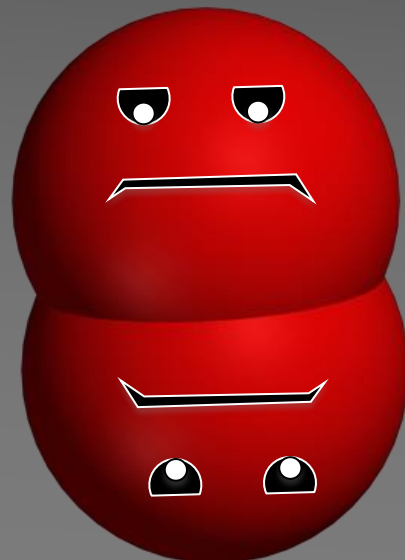


/Combustion/

Everyone loves O₂



+



/Image Credits/

- Title Slide
 - [Aluminum & Copper II Chloride](#) by Mr. W